

A Simple, Rapid Method for the Preparation of [^{11}C]Formaldehyde**

Jacob M. Hooker,* Matthias Schönberger, Hanno Schieferstein, and Joanna S. Fowler

The incorporation of carbon-11 into small molecules has been paramount to the success of positron emission tomography (PET) for in vivo molecular imaging and drug research and development.^[1] However, many of the properties that make ^{11}C an ideal radionuclide for PET have impeded its chemical development. For instance, the short half-life ($t_{1/2} = 20.4$ min), which allows for repeated studies within a short time span, necessitates rapid chemical syntheses and purifications. Moreover, high specific activity, which makes it possible to image low-concentration receptors and molecular targets, places the working concentration range of ^{11}C -labeling reagents in the low nanomolar range. But perhaps the biggest challenge in the synthesis of ^{11}C -labeled compounds is the lack of available labeling reagents. Bear in mind, nearly all carbon-11 syntheses begin with a nuclear reaction [^{14}N -(p, α) ^{11}C] using a cyclotron that produces $^{11}\text{CO}_2$ or $^{11}\text{CH}_4$, from which labeling reagents must be prepared.

By far the most common, almost canonical method, to label a molecule with ^{11}C is through methylation, typically with $^{11}\text{CH}_3\text{I}$.^[2] While pendant methyl groups appear quite frequently in relevant compounds and ^{11}C methylation has led to many successful radiotracers, reliance on methylation limits the range of potential probes. Consequently, new reaction development should focus on methods to incorporate ^{11}C in skeletal positions of target molecules, and several research groups have developed or adapted synthetic methods for ^{11}C incorporation into benzene rings, carbocycles, and heterocycles as well as nonpendant locations.^[3] By using carefully designed organic reactions, each of these has expanded the types of radiotracers that can be accessed.

[^{11}C]Formaldehyde has shown great promise as a labeling reagent for the preparation of PET compounds. Owing to its versatile oxidation state, [^{11}C]formaldehyde has been used in

the syntheses of compounds otherwise unlabelable, through reductive methylations,^[4] ring-closure reactions,^[5] and electrophilic aromatic substitutions,^[6] among others.^[7] However, the widespread development and use of synthetic methods for [^{11}C]formaldehyde in the preparation of PET compounds has been hindered by its limited access. Several methods have been developed for the synthesis of [^{11}C]formaldehyde from [^{11}C]methanol beginning in 1972 here at Brookhaven National Laboratory,^[8] improved over time with new catalysts,^[9] and quite elegantly synthesized enzymatically.^[10] While each of these methods has found utility, they each have disadvantages preventing more prevalent use.

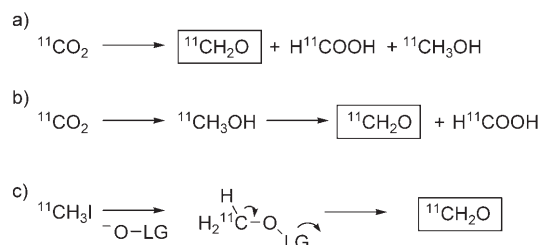
Recognizing the power of [^{11}C]formaldehyde as a labeling reagent, demonstrated in these previous reports, we sought to develop a very simple and rapid method that would provide access to [^{11}C]formaldehyde without the need for any new equipment. To make [^{11}C]formaldehyde instantly available to all chemists interested in isotopically labeled compounds, we also constrained our method to rely on only commercially available starting materials, mild conditions, and short reaction times. Herein, we report a high-yielding method for the production of [^{11}C]formaldehyde that meets all of these criteria.

Previous methods for the preparation of [^{11}C]formaldehyde have relied on the partial reduction of $^{11}\text{CO}_2$ or the complete reduction of $^{11}\text{CO}_2$ to $^{11}\text{CH}_3\text{OH}$ followed by oxidation (Scheme 1). We surmised that using [^{11}C]methyl iodide to access [^{11}C]formaldehyde would be advantageous to these methods for several reasons. First, $^{11}\text{CH}_3\text{I}$ is routinely produced at virtually every location where ^{11}C -labeled compounds are synthesized. Second, we anticipated we could capitalize on the incredible amount of effort that has gone into the development of methods and equipment (now commercially available) for the gas-phase synthesis for high-specific-activity $^{11}\text{CH}_3\text{I}$. Each of the existing methods for [^{11}C]formaldehyde relies on a reduction step that occurs in solution, typically with lithium aluminum hydride, which often causes a reduction in specific activity avoided by gas-phase production of $^{11}\text{CH}_3\text{I}$.^[2] Thus, our efforts began with

[*] Dr. J. M. Hooker, Prof. Dr. J. S. Fowler
Medical Department
Brookhaven National Laboratory, Upton, NY 11973-5000 (USA)
Fax: (+1) 631-344-5815
E-mail: hooker@bnl.gov
M. Schönberger, H. Schieferstein
Johannes-Gutenberg-Universität Mainz (Germany)
Prof. Dr. J. S. Fowler
State University of New York at Stony Brook (USA)
Prof. Dr. J. S. Fowler
Mount Sinai School of Medicine, New York (USA)

[**] This work was carried out at Brookhaven National Laboratory (contract DE-AC02-98CH10886 with the U.S. Department of Energy and supported by its Office of Biological and Environmental Research). J.M.H. was supported by the NIH (1F32EB008320-01) and a Goldhaber Fellowship at BNL. M.S. and H.S. were supported by Deutscher Akademischer Austauschdienst (DAAD) and BNL.

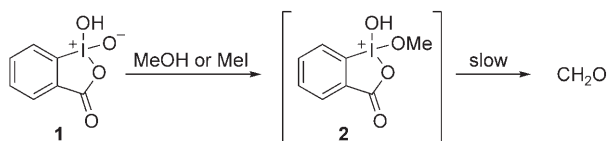
Supporting information for this article is available on the WWW under <http://dx.doi.org/10.1002/anie.200800991>.



Scheme 1. Methods for the production of [^{11}C]formaldehyde: a) by partial reduction of $^{11}\text{CO}_2$; b) by reduction of $^{11}\text{CO}_2$ and subsequent oxidation; c) from [^{11}C]methyl iodide (reported here).

the premise of a mechanistically simple oxidation (Scheme 1c).

Using ^1H NMR spectrometry, we screened the reaction of substoichiometric amounts of methyl iodide with a number of oxygen nucleophiles containing a pendant leaving group (LG). Many inorganic compounds were effective but led to over oxidation. Our first lead resulting exclusively in formaldehyde was found by comparing the reaction of *o*-iodoxybenzoic acid (IBX, **1**) with methanol and methyl iodide (Scheme 2). The common intermediate **2** was observed, which converted to formaldehyde over the course of several days.^[11] It was evident from this study that oxidation rather than methylation was rate limiting. To improve the reaction rate, we turned to trialkylamine *N*-oxides, which we predicted would provide a better leaving group.



Scheme 2. Oxidation of methanol and methyl iodide with IBX (**1**).

By reacting trimethylamine *N*-oxide (TMAO, **3**) with methyl iodide, we observed methylation ($\rightarrow$ **4**) and subsequent formaldehyde formation immediately (as determined by the presence of a ^1H NMR resonance at ca. 9.5 ppm).^[12] In fact, the decomposition of solid **4** to formaldehyde at high temperatures was reported nearly a century ago when the structure of TMAO was under debate.^[13] Given excess

TMAO, **4** underwent elimination to form formaldehyde at room temperature. This process occurred quantitatively at slightly elevated temperatures. Triethylamine *N*-oxide and *N*-methylmorpholine *N*-oxide were also effective but necessitated longer reaction times, whereas pyridine *N*-oxide was ineffective under all condition screened. Upon formation, formaldehyde oligomerized in the presence of TMAO giving rise to multiple methylene resonances at 4.3–4.8 ppm (as is the case in the presence of water).^[14] We used $^{13}\text{CH}_3\text{I}$ to determine if the use of TMAO introduced any isotopic dilution during course of the reaction and to assign all resonances resulting from methyl iodide (Figure 1). We found that the ratio of $^{12}\text{C}/^{13}\text{C}$ remained 1:100 throughout the reaction, suggesting that the methyl groups of TMAO would not impact the specific activity of reactions performed with $^{11}\text{CH}_3\text{I}$.

To test the viability of this method for the preparation of [^{11}C]formaldehyde, we assayed the effect of solvent, TMAO concentration, and reaction time and temperature with the dimedone precipitation method used in previous [^{11}C]formaldehyde reports, (see Table 1). Briefly, an aliquot of the ^{11}C reaction mixture was added to carrier formaldehyde, which was subsequently precipitated as a dimedone adduct. The precipitate was separated by filtration or centrifugation, and the percentage of radioactivity in the precipitate was determined. As a control, we omitted TMAO from the reaction and performed the formaldehyde analysis (for example, see Table 1, entry 5) and found <5% of the radioactivity was associated with the precipitate, which could be removed with washing. We found that DMF was an excellent solvent for the transformation, providing high yields of [^{11}C]formaldehyde in under 2 min at only 70 °C.

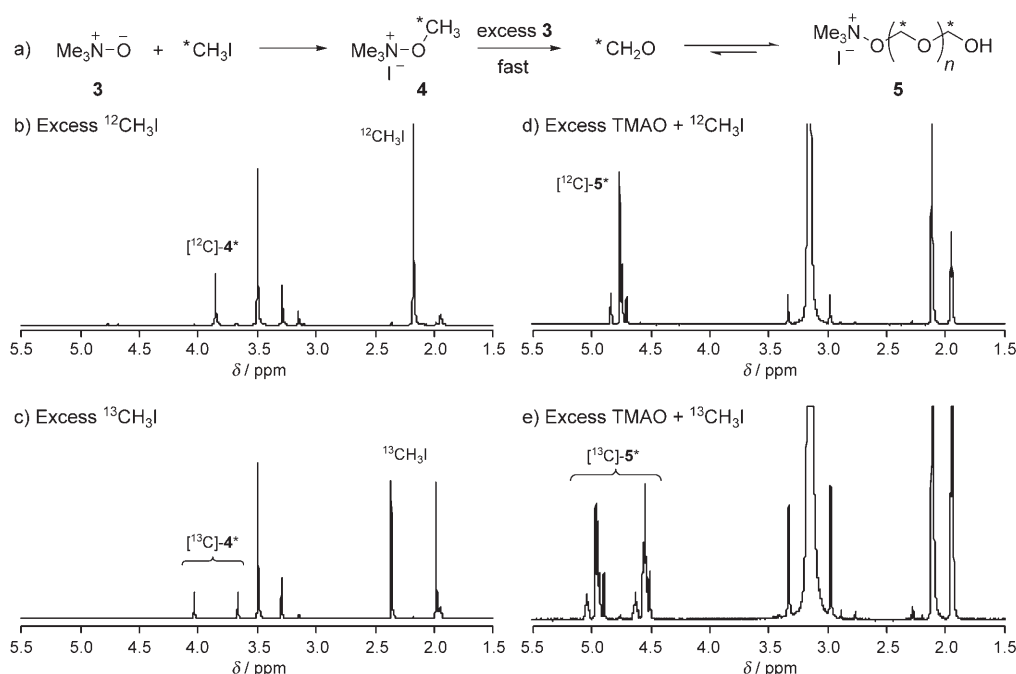


Figure 1. a) The reaction between TMAO and MeI. b–e) ^1H NMR spectra (400 MHz, CD_3CN): b,c) When excess MeI was used, intermediate **4** was observed. d,e) In the presence of excess TMAO, an equilibrium mixture of formaldehyde and its oligomers **5** was observed. Full ^1H NMR spectra as well as ^{13}C NMR spectra are available in the Supporting Information.

Table 1: Optimization of [^{11}C]formaldehyde yield.^[a]

Entry	Solvent	<i>t</i> [s]	TMAO [mg]	<i>T</i> [°C]	Yield [%] ^[b]
1	MeCN	120	4.0	70	27 ± 7
2	MeCN	120	25.0	70	58 ± 8
3	DMSO	120	4.0	70	23 ± 7
4	THF	120	4.0	70	26 ± 4
5	DMF	120	–	70	3 ± 2 ^[c]
6	DMF	120	1.0	70	80 ± 2
7	DMF	30	4.0	70	72 ± 5
8	DMF	60	4.0	70	86 ± 4
9 ^[d]	DMF	120	4.0	70	89 ± 4
10	DMF	120	4.0	20	38 ± 6

[a] Reactions were performed in 300 μL of solvent with ca. 1 mCi $^{11}\text{CH}_3\text{I}$.
[b] Average $\pm$ standard deviation ($n=3$) based on dimedone precipitation method. [c] Radioactivity lost from precipitate through exhaustive washing. [d] Recommended conditions ($n=12$).

We found that the reaction was reproducible and robust. The reaction was not sensitive to small amounts of water (up to 10 μL), and the dihydrate form of TMAO, which is less expensive, was an equally effective reagent. Added base was tolerated as long as it did not competitively methylate. However, the addition of acid lowered reaction yields or prevented [^{11}C]formaldehyde formation altogether. Higher temperatures were equally effective with no overoxidation observed.

In ^{12}C test experiments with nonlabeled compounds, we successfully used the formaldehyde–oligomer mixture **5** in a variety of reaction scenarios including reductive amination of anilines, oxime and hydrazone formation, and cyclization reactions. To highlight the efficacy of this mixture as no carrier added [^{11}C]formaldehyde and to determine specific activity, we performed a Pictet–Spengler condensation with tryptamine (Scheme 3). The reaction of tryptamine with **6** occurred readily under acidic conditions affording [^{11}C]-2,3,4,9-tetrahydro-1*H*- β -carboline (**7**) in excellent radiochemical yield. The product was isolated by semipreparative HPLC using a C18 column, and the mass was determined by absorbance at 254 nm. Using a portion of [^{11}C]MeI from a clinical production run, we determined **7** had a specific activity of 3.0 Ci μmol^{-1} (as compared to 4.5 Ci μmol^{-1} for the clinical product).^[15]

In addition, to further verify TMAO itself does not contribute formaldehyde to the reaction solution, we ran two critical negative control reactions under the same prepara-

tive-scale conditions. In the first, we omitted MeI and found no product formation occurred. In the second, TMAO was omitted, and no [^{11}C]-**7** was observed.

In summary, we have developed a simple and accessible method for production of isotopically labelled formaldehyde. We demonstrated that commercially available, inexpensive trimethylamine *N*-oxide is highly effective at converting [^{11}C]methyl iodide to [^{11}C]formaldehyde under mild conditions with no erosion of specific activity. We anticipate easy access to [^{11}C]formaldehyde will result in the increased use and scope of reactions such as electrophilic aromatic substitution, Mannich-type condensations, and cyclization reactions for the preparation of [^{11}C]-labeled compounds. Current studies are underway to expand the scope and utility of [^{11}C]formaldehyde as a labeling reagent for the synthesis of PET radiotracers.

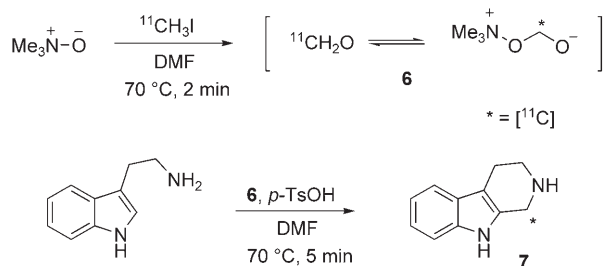
Experimental Section

General procedure for the conversion of $^{11}\text{CH}_3\text{I}$ to $^{11}\text{CH}_2\text{O}$: A mixture of trimethylamine *N*-oxide or its hydrate (4 mg) and DMF (300 μL) cooled to -40°C was used to capture $^{11}\text{CH}_3\text{I}$ produced using PETtrace MeI Microlab (GE Medical Systems, Milwaukee, WI, USA). The sealed vessel was heated to 70°C for 2 min and then cooled in an ice bath.

Received: February 29, 2008

Published online: July 4, 2008

Keywords: formaldehyde · isotopic labeling · *N*-oxides · oxidation · positron emission tomography



Scheme 3. Synthesis of radioactive **7** from tryptamine (**6**) and $^{11}\text{CH}_2\text{O}$. Ts = toluenesulfonyl.

- [1] M. E. Phelps, *Proc. Natl. Acad. Sci. USA* **2000**, *97*, 9226.
- [2] a) B. Långström, H. Lundqvist, *Int. J. Appl. Radiat. Isot.* **1976**, *27*, 357; b) P. Larsen, J. Ulin, K. Dahlström, M. Jensen, *Appl. Radiat. Isot.* **1997**, *48*, 153.
- [3] a) Reviewed in: F. R. Wuest, *Trends Org. Chem.* **2003**, *10*, 61; b) O. Itsenko, D. Norberg, T. Rasmussen, B. Långström, C. Chatgililoglu, *J. Am. Chem. Soc.* **2007**, *129*, 9020, and references cited therein.
- [4] a) M. G. Straatmann, M. J. Welch, *J. Nucl. Med.* **1975**, *16*, 425; b) C. Marazano, M. Mazière, G. Berger, D. Comar, *Int. J. Appl. Radiat. Isot.* **1977**, *28*, 49; c) M. Mazière, G. Berger, D. Comar, *J. Labelled Compd. Radiopharm.* **1977**, *13*, 196; d) G. Berger, M. Mazière, C. Marazano, D. Comar, *Eur. J. Nucl. Med.* **1978**, *3*, 101; e) G. Berger, M. Mazière, R. Knipper, C. Prenant, D. Comar, *Int. J. Appl. Radiat. Isot.* **1979**, *30*, 393.
- [5] a) M. W. Nader, S. K. Zeisler, A. Theobald, F. Oberdorfer, *Appl. Radiat. Isot.* **1998**, *49*, 1599; b) D. Roeda, H. T. Sipila, Y. Bramouille, J. D. Enas, F. Vaufrey, F. Dollé, C. Crouzel, *J. Labelled Compd. Radiopharm.* **2002**, *45*, 37; c) M. van der Mey, A. D. Windhorst, R. P. Klok, J. D. M. Herscheid, L. E. Kennis, F. Bischoff, M. Bakker, X. Langlois, L. Heylen, M. Jurzak, J. E. Leysen, *Bioorg. Med. Chem.* **2006**, *14*, 4526.
- [6] O. Langer, A. Krcal, A. Schmid, A. Abraham, P. Minetti, D. Celona, D. Roeda, F. Dollé, K. Kletter, M. Mueller, *J. Labelled Compd. Radiopharm.* **2005**, *48*, 577.
- [7] V. W. Pike, A. J. Palmer, P. L. Horlock, T. J. Perun, ; L. A. Freiberg, D. A. Dunnigan, R. H. Liss, *Int. J. Appl. Radiat. Isot.* **1984**, *35*, 103; L. A. Freiberg, D. A. Dunnigan, R. H. Liss, *Int. J. Appl. Radiat. Isot.* **1984**, *35*, 103.
- [8] D. Christman, E. J. Crawford, M. Friedkin, A. P. Wolf, *Proc. Natl. Acad. Sci. USA* **1972**, *69*, 988.

- [9] D. Roeda, F. Dollé, *J. Labelled Compd. Radiopharm.* **2003**, *46*, 449.
- [10] a) G. Slegers, R. H. D. Lambrecht, T. Vandewalle, L. Meulewaeter, C. Vandecasteele, *J. Nucl. Med.* **1984**, *25*, 338; b) H. Svård, G. Antoni, S. B. Jigerius, G. Zdansky, B. Langström, *J. Labelled Compd. Radiopharm.* **1984**, *21*, 1175; c) J. A. Hughes, M. Jay, *Nucl. Med. Biol.* **1995**, *22*, 105.
- [11] For reaction kinetics of IBX with MeOH by ^1H NMR see: S. De Munari, M. Frigerio, M. Santagostino, *J. Org. Chem.* **1996**, *61*, 9272.
- [12] H. Dahn, P. Pechy, *Magn. Reson. Chem.* **1996**, *34*, 723.
- [13] J. Meisenheimer, *Justus Liebigs Ann. Chem.* **1913**, 397, 273.
- [14] D. J. Le Botlan, B. G. Mechin, G. J. Martin, *Anal. Chem.* **1983**, *55*, 587.
- [15] “Environmental” formaldehyde is found in many chemicals and solvents and should be rigorously excluded for very high-specific-activity syntheses.